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CONSTANT Calculate the molar gas constant R in the following units: a. $(\text{atm})(\text{cm}^3)/(\text{g}\cdot\text{mol})(\text{K})$ b.

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Source: Otto Frank, "Shortcuts for Distillation Design," Chemical Engineering, March 14, 1977. In the present case, hand calculation and shortcut procedures are adequate because the benzene-toluene-xylene system is close to ideal. 2. Establish the equilibrium relationship among the constituents.

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Handbook of chemical engineering calculations/Nicholas P.Chopey, editor. 3rd ed. p. cm. Includes index. ISBN 0-07-136262-2. 1. Chemical engineering Mathematics. I. Chopey, Nicholas P. TP149.H285 2003.

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Calculation Procedure 1. Assume a basis. Assume gas is at standard conditions, that is, 1 g·mol gas at 1 atm (101.3 kPa) pressure and 0 C (273 K, or 492 R), occupying a volume of 22.4 L. 2. Compute the gas constant. Apply suitable conversion factors and obtain the gas constant in various units. Use $PV = RT$; that is, $R = PV/T$. Thus, a.

SECTION 1 PHYSICAL AND CHEMICAL - ICDST

As a chemical engineer, this book rates up there with Perry's on value of information. Chopey does a very nice job of including detailed step-by-step calculations for a multitude of subjects. The many examples are very relevant to the text and provide good insight to the calculation process. The book is very well organized and easy to use.

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